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CH225 Midterm (35%) Aids allowed: Periodic Table

1. (3) Arrange the following elements in the order of increasing electron affinity: Cl, He, K, S, Li

2. (3) Knowing the atomic radii of C (0.77 Å), Si (1.18 Å), Cl (0.99 Å) and Al (1.43 Å), arrange the following single bonds in the order of increasing bond length: C–Si, C–Al, C–Cl, Si–Cl, Al–Cl,

3. (3) Circle all dative bonds in the following species:

4. (4) Determine the formal oxidation states for all atoms in the following molecules and ions:

CaH_2 OS: Ca ____ H____

$SOCl_2$ OS: S____ O____ Cl____

$[H_3Si_2O_7]^{3-}$ OS: H____ Si____ O____

$F_3B\text{-}NH_3$ OS: F____ B____ N____ H____

$Li[AlCl_2H_2]$ OS: Li____ Al____ Cl____ H____

5. (6) Use VSEPR rules to predict the structure and shape of the following molecules. Prepare clear structural drawings showing all atoms, bonds, as well as lone pairs on the central atom (if present). Do NOT show any other electrons. Name the shapes.

$[BeCl_2F_2]^{2-}$ [$ClO_3]^-$ BrF_5

Molecular shape:

UNIT 16 ELECTROCHEMISTRY REVISION (KCU)

OXIDATION NUMBERS, OXIDATION AND REDUCTION

1. Assign oxidation numbers to each element in the following:

a. MnO_2 Mn + 2O = 0 Mn + 2(-2) = 0 Mn + -4 = 0 ∴ Mn = +4 and O = - 2	b. NO_2 N + 2O = 0 N + 2(-2) = 0 N + - 4 = 0 ∴ N = +4 and O = - 2	c. Cl_2 2Cl = 0 ∴ Cl = 0
d. ClO_4^- Cl + 4O = - 1 Cl + 4(-2) = - 1 Cl + - 8 = - 1 ∴ Cl = + 7 and O = - 2	e. SO_4^{2-} S + 4O = - 2 S + 4(-2) = - 2 S + - 8 = - 2 ∴ S = + 6 and O = - 2	f. Fe_2O_3 Reverse cross-over Fe = +3 and O = - 2 or 2Fe + 3O = 0 2Fe + 3(-2) = 0 2Fe + - 6 = 0 ∴ 2Fe = + 6 Fe = + 3 and O = - 2
g. H_3PO_4 3H + P + 4O = 0 3(+1) + P + 4(-2) = 0 +3 + P + - 8 = 0 ∴ P = + 5, H = +1 and O = -2	h. HPO_4^{2-} H + P + 4O = - 2 +1 + P + 4(-2) = - 2 P - 7 = - 2 ∴ P = + 5, H = +1 and O = - 2	i. CH_4 C + 4H = 0 C + 4(+1) = 0 C + 4 = 0 ∴ C = - 4 and O = - 2
j. CO_2 C + 2O = 0 C + 2(-2) = 0 C + - 4 = 0 ∴ C = + 4 and O = - 2	k. H_2O_2 2H + 2O = 0 2(+1) + 2O = 0 +2 + 2O = 0 ∴ 2O = - 2 O = - 1 and H = +1 (An exception to O = - 2)	l. Cu^{2+} Cu = + 2
m. Ag Ag = 0	n. $Cr_2O_7^{2-}$ 2Cr + 8O = -2 2Cr + 8(-2) = - 2 2Cr + - 16 = - 2 ∴ 2Cr = +14 Cr = +7 and O = - 2	o. IO_3^- I + 3O = - 1 I + 3(-2) = - 1 I + - 6 = - 1 ∴ I = + 5 and O = - 2
p. Mn^{2+} Mn = + 2	q. NO_3^- N + 3O = - 1 N + 3(-2) = - 1 N + - 6 = - 1 ∴ N = + 5 and O = - 2	r. SO_3^{2-} S + 3O = - 2 S + 3(-2) = - 2 S + - 6 = - 2 ∴ S = + 4 and O = - 2
s. SO_2 S + 2O = 0 S + 2(-2) = 0 S + - 4 = 0 ∴ S = + 4 and O = - 2	t. NO_2 N + 2O = 0 N + 2(-2) = 0 N + - 4 = 0 ∴ N = + 4 and O = -2	u. NO N + O = 0 N + -2 = 0 ∴ N = + 2 and O = - 2
v. I^- I = - 1	w. Al^{3+} Al = + 3	x. H_2S 2H + S = 0 2(+1) + S = 0 +2 + S = 0 ∴ S = -2 and H = +1

WORKSHEET: Redox reactions

1. **Define** the process of **oxidation** in terms of :

a) transfer (gain or loss) of oxygen ____________

b) transfer (gain or loss) of electrons ___________

c) transfer (gain or loss) of hydrogen ____________

d) decrease or increase in oxidation number ____________

2. **Define** the process of **reduction** in terms of:

a) transfer (gain or loss) of oxygen ____________

b) transfer (gain or loss) of electrons ___________

c) transfer(gain or loss) of hydrogen ____________

d) decrease or increase in oxidation number__________________

3. a) Complete the half equations by balancing it and identify each half reaction as either an **oxidation (O)** or **reduction (R)** reaction.

(i)	Na (s)	→	Na^+ (aq)	(ii)	Mg^{2+}(aq)	→	Mg (s)
(iii)	Fe (s)	→	Fe^{3+}(aq)	(iv)	Cl^- (aq)	→	Cl_2 (g)
v)	K(s)	→	K^+ (aq)	vi)	O_2 (g)	→	O^{2-} (aq)
vii)	F^- (aq)	→	F_2 (g)	viii)	Cu^{2+} (aq)	→	Cu(s)
ix)	S (s)	→	S^{2-} (aq)	x)	H_2 (g)	→	H^+ (aq)

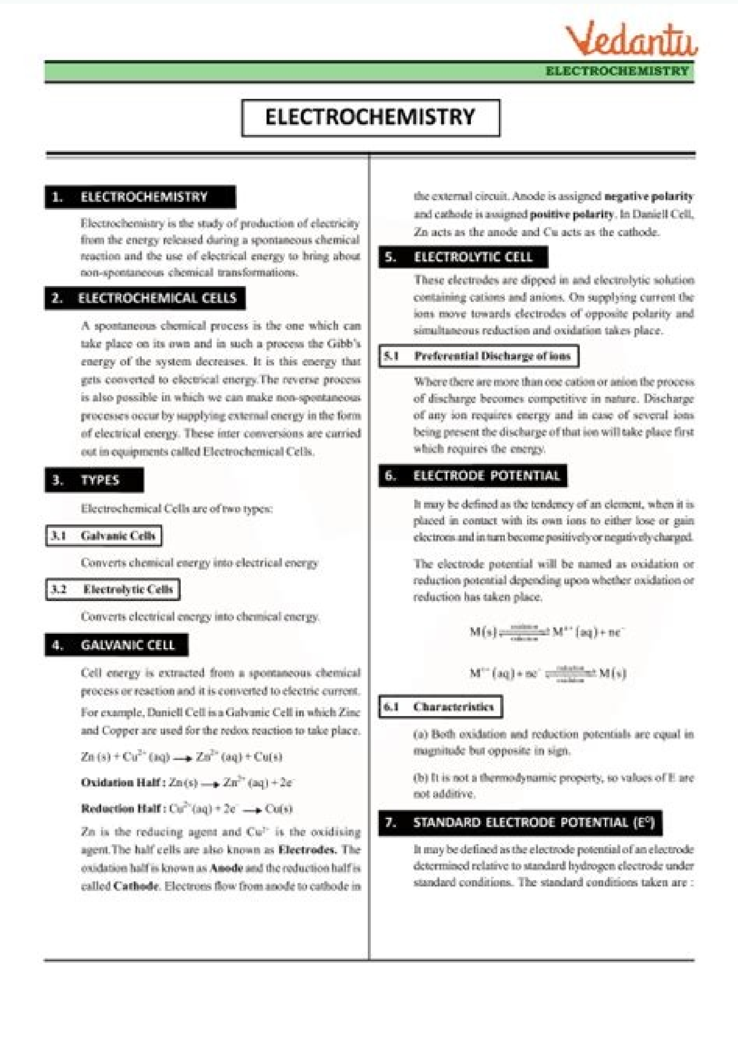
4 Consider the incomplete reaction (partial ionic equation) given below:

A) Cu^{2+} + Zn → Zn^{2+} + Cu

i) Give the two half equations for this reaction; oxidation + reduction

ii) Identify the chemical substance (species) that has undergone oxidation

iii) Identify the chemical substance (species) that has undergone reduction



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Note: Tables 2, 3 and 4 indicate general aging trends of common cobalt-based Li-ion batteries on depth-of-discharge, temperature and charge levels, Table 6 further looks at capacity loss when operating within given and discharge bandwidths. The tables do not address ultra-fast charging and high load discharges that will shorten battery life. No all batteries behave the same.

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